

STANEK, JAROSLAV

Ethyl 3-benzoylacetate. Jaroslav Stanek and Miroslav Holub, *Chem. Pap.*, Prague, Czechoslovakia, 1967, 47, 401-8 (1967). The preparation of  $\text{PhCH:CHCOCH(CO}_2\text{Et)}_2$  (I) and its reactions are described.  $\text{PhCH:CHCOCH(CO}_2\text{Et)}_2$  (I) and 10 g.  $\text{AcCH}_2\text{CO}_2\text{Et}$  heated with 480 ml. 6% aq.  $\text{NaOH}$  yielded 110 g. (73%)  $\text{PhCH:CHCOCH(CO}_2\text{Et)}_2$  (II), m. 11°. II (10 g.) heated with 40 ml. 75%  $\text{AcOH}$  3 hrs. on the steam bath gave 3.5 g. (41%) I, m. 43° (from  $\text{EtOH}$ ). A better yield (60%) of I, m. 43.5°, was obtained by alk. cleavage by shaking 50 g. II 10 min. with 150 ml. 8% aq.  $\text{NaHCO}_3$ . I (1 g.) boiled briefly with 0.33 g.  $\text{H}_2\text{NOH}\cdot\text{HCl}$  in 10 ml.  $\text{EtOH}$ , gave 2-styryl-2,4-oxazolin-5-one, m. 163° (from  $\text{EtOH}$ ). Heating 1 g. I 30 min. on the steam bath with 0.5 g.  $\text{PhNHNH}_2$  gave 0.6 g. 1-phenyl-3-styryl-2-pyrazolin-5-one, m. 150° (from  $\text{EtOH}$ ). I (0.5 g.) boiled 10 min. with 50 ml. 1%  $\text{NaOH}$  soln. of 2,4-( $\text{O}_2\text{N}$ )- $\text{C}_6\text{H}_3\text{NNH}_2$  yielded  $\text{PhCH:CHC[CH(CO}_2\text{Et)}_2\text{NHC}_6\text{H}_3(\text{NO}_2)_2$  (III), m. 152° (from  $\text{EtOH}$ ), and 1-(2,4-dinitrophenyl-3-styryl-2-pyrazolin-5-one, m. 171° (from  $\text{EtOH}$ ). The Na salt of I, prepd. from 5 g. I and 0.52 g. Na in 50 ml.  $\text{Et}_2\text{O}$  with 2.8 g. iodine in 30 ml.  $\text{Et}_2\text{O}$  gave  $\text{PhCH:CHCOCH(CO}_2\text{Et)}_2$ , m. 154° (from  $\text{MeOH}$ ). I (2 g.) heated gently with 0.4 g. 40% aq.  $\text{CH}_2\text{O}$  or with 0.15 g. ( $\text{CH}_2\text{O}$ ), in the presence of 3 drops  $\text{C}_6\text{H}_5\text{N}$  gave  $\text{PhCH:CHCOCH(CO}_2\text{Et)}_2$ , m. 118°. Treatment of 10 g. I in 100 ml.  $\text{Et}_2\text{O}$  5 hrs. with  $\text{NH}_3$  in the presence of 10 g.  $\text{NH}_4\text{NO}_3$  (or treatment of I with  $\text{PhNH}_2$  or  $\text{CH(CO}_2\text{Et)}_2\text{CHPhCH:CHCOCH(CO}_2\text{Et)}_2$  (IV), which, boiled 3 hrs. with equal amt. of  $\text{H}_2\text{NOH}\cdot\text{HCl}$  in  $\text{EtOH}$  yielded 2-carbethoxymethyl-1-phenyl-3-styryl-2-pyrazolin-5-one, m. 148° (from  $\text{CHCl}_3$  and  $\text{EtOH}$ ) (0.7 g. from 2 g. IV).

M. Hudlický

STANEK, J.

Chemical Abst.  
Vol. 48 No. 6  
Mar. 25, 1954  
Organic Chemistry

2,4-Dithio-5,5-diphenylhydantoin. J. Staněk and J. Štílo  
(Charles Univ., Prague, Czech.). *Chem. Listy* 47, 470  
(1953).—Refusing 24 g. 5,5-diphenylhydantoin with 60 g.  
P<sub>2</sub>S<sub>5</sub> in 400 ml. tetralin 2 hrs. gave 18 g. 2,4-dithio-5,5-di-  
phenylhydantoin (I), m. 246° (from EtOH). No desulfura-  
tion of I was observed by the action of Raney Ni.  
M. Hudlický

STANĚK, J.

Chemical Abst.  
Vol. 48 No. 6  
Mar. 25, 1954  
Organic Chemistry

4,6-Dioxohexahydropyrimidine. I. Staněk and J. Štílo  
(Charles Univ., Prague, Czech.). *Chemistry* 47, 471  
(1953).—4,6-Dioxohexahydropyrimidine (yield 12%), m.  
148°, and 1,3-diphenyl-4,6-dioxohexahydropyrimidine (yield  
25%), m. 178°, were prepd. by refluxing, in EtOH with  
Raney Ni, thiobarbituric acid and 1,3-diphenylthiobarbi-  
turic acid, resp. M. Hudlický

Reactivity of the methyl group on the heterocyclic nucleus. IV. J. Stinner and V. Zedlitz (Charles Univ., Prague). *Chem. Listy* 47, 471-2 (1953); cf. *C.A.* 47, 12378i. 2-Phenacyl-6-methylpyridine, m. 75-6° (from petr. ether) (cf. *C.A.* 47, 9972a) was found to be unstable and was changed into its stable isomer form, m. 131° by recrystn. from ether. Nicholas Feldman.

ADJ BZ

STANEK, JAROSLAV

Phenyl glycosides of D-cellobiose. Jaroslav Staněk and Jan Kocourek (Charles Univ., Prague, Czech.). *Chem. Listy* 47, 697-702 (1953).—Octaacetyl- $\alpha$ -D-cellobiose (29) g.) in 400 ml.  $\text{CHCl}_3$  treated 12 hrs. with 450 g. HBr in 760 ml. AcOH, poured onto ice, the aq. layer extd. with  $\text{CHCl}_3$ , and the ext. evapd. to beginning crystn. and mixed with an equal vol. of petr. ether yielded 92.3 g. (74.5%) heptaacetyl- $\alpha$ -D-cellobiosyl bromide (I), m. 185°. I (20 g.) in 60 ml. MeCO allowed to stand 24 hrs. with 10 g. PhOH and 2.5 g. NaOH in 40 ml.  $\text{H}_2\text{O}$ , the solvent distd. *in vacuo*, the residue dild. with an equal vol. of  $\text{H}_2\text{O}$ , extd. with four 25-ml. portions of  $\text{CHCl}_3$ , and the ext. evapd. and dild. with Et<sub>2</sub>O yielded 10 g. Ph heptaacetyl- $\beta$ -D-cellobioside (II), m. 214°, giving 7.5 g. (37%) pure II, m. 216° (from MeOH). Sapon. of 20 g. II with 8 ml. 0.01N MeONa in 50 ml. MeOH gave 7.3 g. (57.4%) Ph  $\beta$ -D-cellobioside, m. 209°; dihydrate, m. 129° (from  $\text{H}_2\text{O}$ ),  $[\alpha]_D^{25} -53.1^\circ$ . I (31.5 g.), 15 g. p-HOC<sub>6</sub>H<sub>4</sub>OBz, 2.5 g. NaOH, 40 ml.  $\text{H}_2\text{O}$ , and 50 ml. MeCO gave 18 g. p-benzoyloxyphenyl heptaacetyl- $\beta$ -D-cellobioside (III), m. 244°,  $[\alpha]_D^{25} -25.3^\circ$ . Sapon. of 9 g. III with 9 ml. 0.1N NaOMe in 70 ml. MeOH gave 3 g. (63.4%) p-hydroxyphenyl  $\beta$ -D-cellobioside (IV), m. 216°,  $[\alpha]_D^{25} -53.0^\circ$ . Acetylation of IV gave 67.7% p-acetoxyphenylheptaacetyl  $\beta$ -D-cellobioside, m. 220° (from EtOH- $\text{CHCl}_3$ ),  $[\alpha]_D^{25} -52.48^\circ$ . I (10 g.) dissolved in 120 ml.  $\text{C}_6\text{H}_6$ , refluxed with 4 g. PhOH and 2 g. Hg(OAc)<sub>2</sub> 1 hr., the mixt. filtered, evapd., heated 2 hrs. *in vacuo* at 60°, and dild. with 120 ml. boiling EtOH, yielded 3.5 g. after crystn. 1.4 g.) Ph heptaacetyl- $\alpha$ -D-cellobioside (V), m. 221° (from EtOH- $\text{CHCl}_3$ ),  $[\alpha]_D^{25} 75.0^\circ$ . The same expt. carried out in the presence of 6 g. Al gave 4.5 g. V, m. 225° (after recrystn.),  $[\alpha]_D^{25} 78.7^\circ$ . The usual sapon. of V with MeONa gave 59.8% Ph  $\alpha$ -D-cellobioside, m. 252-4°,  $[\alpha]_D^{25} 122.91^\circ$ . I (20 g.) and 8 g. PhOH dissolved in 100 ml. MeCO, refluxed 2 hrs. with 4.4 g. Hg(OAc)<sub>2</sub>, cooled, filtered, evapd. to dryness *in vacuo* and the residue treated with 100 ml. Et<sub>2</sub>O yielded 14 g. (10.4 g., 55% after recrystn.) of octaacetyl- $\beta$ -D-cellobiose, m. 192-3°,  $[\alpha]_D^{25} -15.2^\circ$ , also obtained in 82% yield  $[\alpha]_D^{25} -11.58^\circ$  by treatment of I with AgOAc. M. Hudlický

STANEK, J., JAROLIM, V.

"Condensation of Succinic Acid Dinitrile with Alicyclic Ketones" p. 703,  
(CHEMICKÉ LISTY, Vol. 47, no. 5, May 1953, Praha, Czechoslovakia).

SO: Monthly List of East European Accessions, LC, Vol. 2, No. 11, Nov. 1953, Uncl.

SIANEK-1

Reactivity of the methyl group on the heterocyclic ring.  
 V. The base-catalyzed condensation of 2-picoline methiodide with aromatic aldehydes. I. Staneš and Z. Zeska (Charles Univ., Prague, Czech.). *Chem. Listy* 47, 748-51 (1953); cf. *C.A.* 47, 12378i; 48, 3370f. — Nitro- and halo-benzaldehydes in the presence of  $C_6H_5N$  with 2-picoline-MeI form methiodides of 2-(2-aryl-2-hydroxyethyl)pyridines (I) in the cold (Type 1 reaction), and methiodides of substituted 2-stilbazoles (II) at higher temp. (Type 2 reaction). Other aromatic aldehydes react always according to the type 2 reaction. Type 1 (aryl group, % yield, m.p.):  $o-O_2NC_6H_4$ , 38, 140°;  $m-O_2NC_6H_4$ , 60, 137°;  $p-O_2NC_6H_4$ , 39, 133°;  $m-ClC_6H_4$ , 38, 148°;  $p-BrC_6H_4$ , 35, 132°. Type 2 (aryl group, % yield, m.p.):  $o-O_2NC_6H_4$ , 32, 235°;  $m-O_2NC_6H_4$ , 45, 238°;  $p-O_2NC_6H_4$ , 51, 241°;  $m-ClC_6H_4$ , 60, 201°;  $p-BrC_6H_4$ , 72, 205°;  $p-Me_2NC_6H_4$ , 84, 265°;  $o-HOC_6H_4$ , 45, 248°;  $o-EtOC_6H_4$ , 40, 239°; *furyl*, 67, 194°; *Ph*, 20, 188°. M. Hudlický

AA Jan

STANĚK, J.

Chemical Abst.  
Vol. 48  
Apr. 10, 1954  
Organic Chemistry

Reaction of silver salts of carboxylic acids with halogens  
(the Hunsdiecker reaction). Jaroslav Staněk (Charles  
Univ., Prague, Czech.). *Chem. Abstr.* 1954, 48, 124-125 (1953).—A  
review with 65 references and tables of compds.  
M. Hudlický.



STANEK, J.; HAUZAR, I.

"Preparation of 6-Acetyl-1, 2-Isopropylidene- $\alpha$ -D Glucofuranose", P. 337  
(CHEMICKE ZVESTI, Vol. 8, No. 6, June 1954, Bratislava, Czechoslovakia)

SO: Monthly List of East European Accessions, (EEAL), LC, Vol. 3, No. 12,  
Dec. 1954, Uncl.

CZECH

Preparation of 2-deoxy-D-glucose. Jaroslav Staněk and  
~~Oliver Schwarz (Karlova Univ., Prague)~~ ~~Chem. Listy~~  
~~48, 370-81 (1954); Collection Czechoslov. Chem. Commun.~~  
~~20, 42-5 (1955) (in German).~~ Triacetyl-D-glucal (I) added  
 (BzO)<sub>2</sub>Ag and (BzO)<sub>2</sub>BrAg to give 1-benzoyl-3,4,6-tri-  
 acetyl-2-deoxy-2-iodo-α-D-glucopyranose (II) and 1-benzoyl-  
 3,4,6-triacetyl-2-deoxy-2-bromo-α-D-glucopyranose (III), resp.  
 Both halogen derivs. give 2-deoxy-D-glucose (IV) by reduction.  
 Adding a C<sub>6</sub>H<sub>6</sub> soln. of 16.7 g. iodine into a suspension  
 of 23.6 g. dry BzOAg in 200 ml. CCl<sub>4</sub>, treating the mixt.  
 with a soln. of 20 g. I in 200 ml. CCl<sub>4</sub>, heating the mixt.  
 7 hrs. on the steam bath, removing the AgI, evap. the sol-  
 vent, and crystg. the residue from EtOH gave 20.8 g.  
 (54.7%) II, m. 120-30°, [α]<sub>D</sub><sup>20</sup> 21.7°. Analogous procedure  
 with 13.4 g. BzOAg, 4.6 g. Br, and 3 g. I gave 3.9 g. (33%)  
 III, m. 130-40°, [α]<sub>D</sub><sup>20</sup> 33.5°. The same compd. (3 g.), m.  
 140°, [α]<sub>D</sub><sup>20</sup> 33.6°, was obtained by adding 3.3 g. Br to a  
 soln. of 5.44 g. I in 50 ml. CCl<sub>4</sub>; by refluxing the mixt. 2  
 hrs. with 6 g. BzOAg, filtering off the AgBr, and evap.  
 the solvent. Reducing 8 g. II or III in 150 ml. Me-  
 OH with 60 g. Zn activated by 1 hr. immersion in a soln. of  
 60 g. CuSO<sub>4</sub> in 1500 ml. H<sub>2</sub>O, removing Zn after 8 hrs.,  
 evap. the MeOH, and syong. the residue with Ba(OH)<sub>2</sub>  
 yielded 0.43 g. (20%) IV, m. 145°, [α]<sub>D</sub><sup>20</sup> 46.1°.

M. Hudlický

STANEK, Jaroslav

✓ The identity of bellamarine and acetylcaranine. Jaroslav Stanek (Charles Univ., Prague). *Chemistry & Industry* 1955, 1567. — The identity of bellamarine (I) was confirmed as being acetylcaranine. Pure I from chromatography m. 185°,  $[\alpha]_D^{25}$  -183.5, and contains the OCH<sub>3</sub>O group.  
C. Robert Walter, Jr.

STANEK, J.; SCHWARZ, V.

Preparation of 2-desoxy-D-glucose. In German. p. 42

Vol. 20, no. 1, Feb. 1955  
SBORNIK CHEKHOSLOVATSKIKH KHMICHESKIKH RABOT  
Praha, Czechoslovakia

So: Eastern European Accession Vol. 5, No. 4, 1956

Stanek, J.

✓ The anomeric configuration of hemiacetal groups in glycoalkaloids. J. Stanek (Charles Univ., Prague). *Chemistry & Industry*, 1967, 138-140. The author has applied Klyne's rule for the calcn. of the mol. rotation of glucosides contg. an optically active aglycon to the calcd. and found mol. rotations for a series of glycoalkaloids and observed that the agreement is generally good. Some of the hydrolysis products are not true aglycons and the calcons. are of little value. The rule does not hold for some synthetic glycosides such as those of the morphine group. Anomalous rotations

may occur due to hydrolytic isomerizations. Tables showing the mol. rotations of some glycoalkaloids and some anomalies are listed.

Gilles Flower, Jr.

mm

STANEK, JAROSLAV

Alkaloidy. Praha, Maki. Ceskoslovenske akademie ved. 1957. 653 p. (Ceskoslovenska akademie ved. Sekce chemicka. Studie a prameny, sv. 25) (Alkaloids. 1st ed. bibl., footnoest, indexes)

SO: Monthly Index of East European Accessions (EEA) LC, Vol. 7, No. 1, Jan 1958

STANEK, J.

CZECHOSLOVAKIA/Chemical Technology. Chemical Products and Their  
Applications. Pharmaceuticals. Vitamins. Antibiotics.

H

Abs Jour: Ref Zhur-Khim., No 8, 1959, 28581.

Author : Stanek, J.

Inst :

Title : On the Nomenclature Used for Organic Compounds in  
the Czech Pharmacopoeia. Communication 2.

Orig Pub: Ceskoslov Farmac, 6, No 5, 270-272 (1957) (in Czech)

Abstract: No abstract.

Card : 1/1

STANEK, J.

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G-2

Abs Jour: Referat Zhur-Khimiya, No 4, 1958, 11246.

Author : Stanek, J. and Zemlicka, J.

Inst :

Title : Oxidation of Several  $\alpha$ -Substituted Alcohol Derivatives by the Oppenauer Method.

Orig Pub: Chem Listy, 51, No 3, 493-496 (1957) (in Czech)

Abstract: A number of  $\alpha$ -ketols and 1,2-glycols have been oxidized to the corresponding  $\alpha$ -diketones by a modified Oppenauer procedure using Al-phenoxide (I) and p-benzoquinone (II) in anhydrous media (see M. Yamashita and T. Matsumura, J. Chem. Soc., Japan, 64, 506 (1943); R. L. McKee and E. R. Henze, J. Amer. Chem. Soc., 66, 2021 (1944); P. Ruggli et al, Helv. Chim. Acta, 29, 312 (1946)). The esters of  $\alpha$ -hydroxy acids and of  $\alpha$ -amino alcohols

Card : 1/3

7

APPROVED FOR RELEASE: 08/25/2000 CIA-RDP86-00513R001652820011-5"

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry.

Abs Jour: Referat Zhur-Khimiya, No 4, 1958, 11246.

do not react under the above conditions. Preparation: 0.1 mol of the substance to be oxidized is heated 4 hrs in 350 ml  $C_6H_6$  with 0.05 mol I and 0.1-0.15 mol II (per hydroxyl group to be oxidized); the solution is filtered and extracted four times with 200 ml 5% NaOH and water, the  $C_6H_6$  is distilled off, and the residue is distilled or recrystallized. The substance to be oxidized, the oxidation product, and the yield of the latter in % are given below: hydrobenzoin, benzil, 74; iso-hydrobenzoin, benzil, 75; benzoin, benzil, 87; furoin, furil, 49; 1-phenyl-2-nitroethanol, resin, -; D-borneol, D-camphor, 76.5; benzyl alcohol, benzaldehyde, 31.1; cyclohexanol, cyclohexanone, 29.8 (the reaction is accompanied by the self-condensation of cyclohexanone to yield 2-cyclohexenyl-cyclohexanone, bp 108-110°/3 mm, 12.4%). Acetone glycerine, 1,2- and

Card : 2/3



STANEK, J., AND OTHERS

"Reaction of 2, 3, 4, 6-tetraacetyl- $\alpha$ -D-glucopyranosyl bromide with mercaptans."  
p.1556 (Chemicke Listy, Vol. 51, no. 8, Aug. 1957, Praha, Czechoslovakia.)

Monthly Index of East European Accessions (EEAI) LC, Vol. 7, No.6 June 1958.

STANEK, JAROSLAV

Optical rotation of the isomeric trehaloses. Jaroslav Stanek (Charles Univ., Prague). *Nature* 179, 97-8 (1957).  
 The hypothesis that the mol. rotation of any glycoside contg. an optically active aglycone is equal to the sum of the mol. rotations of the corresponding methylglycoside and the aglycone does not apply to the mol. rotations of disaccharides. The calcd. values are:  $\alpha$ -D-glucopyranosyl- $\alpha$ -D-glucopyranose,  $\alpha$ , $\alpha$ -trehalose,  $[\alpha]_D + 61.720^\circ$ ,  $[\alpha]_D + 180.5^\circ$ , octaacetate  $[\alpha]_D + 91.560^\circ$ ,  $[\alpha]_D + 139.3^\circ$ ;  $\alpha$ -D-glucopyranosyl- $\beta$ -D-glucopyranose,  $\alpha$ , $\beta$ -trehalose (neotrehalose),  $[\alpha]_D + 24.220^\circ$ ,  $[\alpha]_D + 70.8^\circ$ , octaacetate  $[\alpha]_D + 40.690^\circ$ ,  $[\alpha]_D + 59.9^\circ$ ;  $\beta$ -D-glucopyranosyl- $\beta$ -D-glucopyranose,  $\beta$ , $\beta$ -trehalose (isotrehalose)  $[\alpha]_D - 13.280^\circ$ ,  $[\alpha]_D - 38.5^\circ$ , octaacetate  $[\alpha]_D - 13.180^\circ$ ,  $[\alpha]_D - 19.4^\circ$ .  
 F. J. Vara

Change 1

CZECHOSLOVAKIA / Organic Chemistry. Natural Substances and Synthetic Analogues. G

Abs Jour: Ref Zhur-Khimiya, No 18. 1958, 61053.

Author : J. Stanek, K. Malkovsky, M. Novak, D. Petricek.

Inst : -

Title : Interaction of 2,3,4,6-Tetraacetyl- $\alpha$ -D-Glucopyranosylbromide with Mercaptans.

Orig Pub: Collect. czechosl. chem. commun., 1958, 23, No 2, 336-338.

Abstract: See RZhKhim, 1958, 39741.

Card 1/1

48

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CIA-RDP86-00513R001652820011-5"

Country : CZECHOSLOVAKIA

G

Category : Organic Chemistry. Synthetic Organic Chemistry

Abs. Jour : Ref Zhur - Khim., No 5, 1959, No. 15430

Author : Stanek, J.

Institut. : -

Title : On the Formation of 2-Alkylmercapto Derivatives of 4-Methyl-6-Oxypyrimidine

Orig Pub. : Collect. czechosl. chem. commun., 1958, 23, No 6, 1154-1156

Abstract : No abstract.

See Ref Zhur-Khim, 1959, 1304.

Card: 1/1

CZECHOSLOVAKIA/Organic Chemistry. Natural Products and Their  
Synthetic Analogues

G-3

Abs Jour: Ref Zhur-Khim, No 24, 1958, 81741.

reaction with aniline (IV) proceeds slowly and with p-nitroaniline does not usually take place. The reaction of 2,3,4,6-tetraacetyl- $\alpha$ -D-glucopyranose with IV takes place very easily, with the formation of pure N-phenyl-2,3,4,6-tetraacetyl- $\alpha$ -D-glucosylamine. The reaction of II with o- and p-III gives pure  $\alpha$ -anomer whereas the same reaction with IV produced a mixture of both anomers of the pure  $\alpha$ -anomer is formed in a low yield (after long boiling). The 2,3,4,6-tetraacetyl- $\beta$ -D-pyranosyl mercaptan (V) reacts in the same way as II but somewhat slower. To obtain the acylated substitutes, three methods are used: A) 0.01 moles of I and 0.02 moles of aromatic

Card : 2/5

CZECHOSLOVAKIA/Organic Chemistry Natural Products and Their  
Synthetic Analogues.

G-3

Abs Jour: Ref Zhur-Khin., No 24. 1958, 81741.

in a 42% yield, m.p. 97°C.  $[\alpha]_D^{25}$  -56.6° (chloro-  
form), with method B, in a 56% yield,  $[\alpha]_D^{25}$  -57.1°,  
by method B some formation of the  $\alpha$ -anomer is ob-  
served. The  $\alpha$ -anomer is obtained with method B,  
yield 39%, m.p. 148°C.  $[\alpha]_D^{25}$  +172.8°. Further  
is given the product, the method of preparation, the  
yield in %, m.p. in °C.  $[\alpha]_D^{25}$  (in chloroform):  
N-phenyl tetraacetyl- $\beta$ -D-galactopyranosylamine,  
A, 43, 119, -32.3°; N-o-tolyltetraacetyl- $\beta$ -D-  
glucopyranosylamine, A, 36, 109, -62.6°; CB, 42,  
109, -63.8°; N-p-tolyl tetraacetyl- $\beta$ -D-glucopyrano-  
sylamine, A, 52, 146, -36.3°; B, 58, 147, -36.8°; C,  
72, 148, -34.4°; N-o-tolyl tetraacetyl- $\beta$ -D-

Card : 4/5

Card : 5/5

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G

Abs Jour: Ref Zhur-Khin., No 2, 1959, 4619.

CuO, are cyclized by the action of strong acids in  $\text{CH}_3\text{COOH}$  medium to form unsaturated 2,5-disubstituted derivatives of furan of the type  $\text{OCR}=\text{CHCH}=\text{CCH}=\text{CHR}'$ . The reaction in all probability proceeds by a mechanism similar to that of the opening of the furan ring according to Marchwald. Preparation: 60 gms of butyl chloride on treatment with diazomethane in ether solution at  $-20^\circ$  give 1-diazopentanone-2; the ether is distilled off and the product is decomposed by refluxing for 15 min with 6 gms CuO in 2 liters of  $\text{C}_6\text{H}_6$ , giving 5-decene-4,7-dione (I), yield 28.5%, mp  $55-56.5^\circ$  (from  $\text{CH}_3\text{OH}$ ). Using a similar procedure, dihydrocinnamyl chloride gives a 27.3% yield of 1,8-diphenyl-4-octene-3,6-dione (II), mp  $85-85.5^\circ$  (from alc). 2 gms of the methyl

Card : 2/9

8

APPROVED FOR RELEASE: 08/25/2000

CIA-RDP86-00513R001652820011-5"

... of the solvent gives 1.85 gm of 5-(3-carbonethoxyethyl)-2-(3-carbonethoxyvinyl)-furan, mp  $59-60^\circ$  (from  $\text{CH}_3\text{OH}$  at  $-30^\circ$ ); refluxing for 30 min with 10% soda solution gives the free acid (III), mp  $178-179^\circ$  (from  $\text{CH}_3\text{OH}$ ). When it is stand in a 0.065 N HCl solution in 99.8%  $\text{CH}_3\text{COOH}$  (90 min  $25^\circ$ ), a 76% yield of 5-phenetyl-2-styryl furan is obtained, mp  $51-51.5^\circ$ . Using a similar procedure, the diethyl ester of eicosene-10-dione-9,12-dicarboxylic-1,20 acid is made to give 65% yields of 5-(8-carbethoxy-octyl)-2-(8-carbethoxy-1-octenyl)-furan; the product is purified by

Card : 3/9

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G

Abs Jour: Ref Zhur-Khin., No 2, 1959, 4619.

chromatography on  $\text{Al}_2\text{O}_3$  (elution with benzene); the free acid (Mp  $101^\circ$ ; from aqueous alcohol) on oxidation with  $\text{KMnO}_4$  gives suberic acid and sebacic acid. I yields 2-propenyl-5-propylfuran, bp  $47-50^\circ/0.4$  mm,  $n_D^{20}$  1.5008; the methyl ester of 4-heptene-3,6-dione-1-carboxylic acid gives 5-methyl-2-(5-carbonethoxyvinyl)-furan (IV), yield 74%, mp  $55-56^\circ$ ; the free acid (V) has an mp of

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G

Abs Jour: Ref Zhur-Khim., No 2, 1959, 4619.

report were obtained by the decomposition of the diazoketones in the presence of the corresponding esters of diazoketocarboxylic acids. In addition to the symmetric product the reaction always yields two symmetric compounds. When 0.05 mol of the methylester of  $\omega$ -diazolevulinic acid (VII) and 0.05 mol of diazoacetone (VIII) are refluxed for 15 min with 3 gms CuO in 1 liter  $C_6H_6$ , 22% yields of the methyl ester of 4-heptene-3,6-dione-1-carboxylic acid are obtained, bp 119-120°/2mm; crystallization gives the dimethyl ester of 4-octene-3,6-dione-1,8-dicarboxylic acid, yield 19%, mp 120-121° (from  $CH_3OH$ ), and distillation gives diacetylene (IX), yield 36%, bp 80-85°/14mm, mp 75° (from ether). When a three-

Card : 5/9

ner Zhur-Khim., No 2, 1959, 4619. G

fold excess of VIII is used, the yield [sic] is increased to 33%. Using a similar procedure, VII and 1-diazo-2-butanone give the methyl ester of 4-octene-3,6-dione-1-carboxylic acid, yield 15.5%, bp 125-130°/2mm, mp 47-48° (from petroleum ether) (the product was separated by distillation after the removal of 1,2-dipropionyl ethylene, yield 53%, bp 80-85°/3mm, mp 52-53° (from petroleum ether)); VIII and the methyl ester of  $\omega$ -diazoacetylvalerianic acid after distillation of IX (47%) and crystallization of the methyl ester of dodecene-6-5,8-dione-1,12-dicarboxylic acid (yield 14%, mp 93-94° (from  $CH_3OH$ )) give the methyl ester of 6-nonene-5,8-dione-1-carboxylic acid, yield 21%, bp 120-140°/2 mm, mp 53-54° (from petroleum ether).

Card : 6/9

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G

Abs Jour: Ref Zhur-Khim., No 2, 1959, 4619.

Abs Jour: Ref Zhur-Khim., No 2, 1959, 4619.

benzoic acid (X), yield 98%, mp 70-71°; the ethyl ester of p-diazoacetylbenzoic acid (XI) yield 99%, mp 65-66°. The following compounds were prepared by refluxing X or XI for 15 min with CuO in C<sub>6</sub>H<sub>6</sub>: 33.5% 1,2-bis-(m-carbethoxybenzoyl)-ethylene, mp 131-132° (from ethyl acetate) (XII); 36.5% 1,2-bis-(p-carbethoxybenzoyl)-ethylene (XIII), mp 194-195° (from benzene). 1,2-bis-(o-carbethoxybenzoyl)-ethylene, mp 160° (decomp; from alc-benzene) was obtained by a similar procedure from the acid ethyl ester of phthalic acid without the separation of intermediate products. The hydrogenation of XIII to XII over Pt (from PtO<sub>2</sub>) at 22° and at normal pressure yields 1,2-bis-(p-carbethoxylbenzoyl)-ethane, mp 157-158° (from alc); similarly 1,4-bis-

Card : 8/9

11

Abs Jour: Ref Zhur-Khim., No 24, 1958, 81743.

Author : Cerny M., Vrkoc J., Stanek J.

Inst :

Title : The Problem of Preparing Acylated Derivatives of Glycopyranosyl Mercaptans

Orig Pub: Chem. listy, 1958, 52, No 2, 311-315.

Abstract: When 333 grams of 2,3,4,6-tetraacetyl- $\alpha$ -D-glycopyranosyl bromide was heated for 15 minutes to boiling point with 60 grams of thiourea in 300 ml acetone and after cooling with ice one obtains 300 grams of crude 2,3,4,6-tetraacetyl- $\alpha$ -D-glucopyranosyl-isothiuronium bromide (I), m.p. 178°C., applied for further investigation. The pure I, m.p. 189°C. (from

Card : 1/6



CZECHOSLOVAKIA/Organic Chemistry. Natural Products and Their  
Synthetic Analogues.

G-3

Abs Jour: Ref Zhur-Khim., No 24, 1958, 81743.

solution of potassium carbonate and allowing it to stand for 30 minutes, yield 82%. III is dimorphic, its modification with a m.p. of 115°C. (from ether or methanol) is obtained by allowing an aqueous solution of I to stand for 30 minutes with a saturated solution of NaHSO<sub>3</sub> or by boiling it for 5 minutes, yield 75 or 85%. The oxidation of both forms of III with 10% H<sub>2</sub>O<sub>2</sub> (one hour at 20°C) in methanol and standing in the refrigerator resulted in the formation of 2,3,4,6,2',3',4',6'-octaacetyl-β-D-glucopyranosyl disulfide (IV), yield 90%, m.p. 143-144°C. (from CH<sub>3</sub>OH), [α]<sub>D</sub><sup>20</sup> -156.5 ± 0.7° (c 2; CHCl<sub>3</sub>), [α]<sub>D</sub><sup>20</sup> -160.3 ± 1.7° (c 0.6; chloroform). IV was also obtained by heating an aqueous suspension of II to boil-

Card : 3/6

APPROVED FOR RELEASE: 08/25/2000  
Synthetic Analogues.

CIA-RDP86-00513R001652820011-5"

G-3

Abs Jour: Ref Zhur-Khim., No 24, 1958, 81743.

ing, then dissolving the resulting syrup in methanol and concentrating by boiling, yield 25%. After both forms of III were allowed to stand for 12 hours with acetic anhydride in pyridine, 1-S-acetyl-2,3,4,6-tetra-O-acetyl-β-D-glucopyranosyl mercaptan (V) was formed, yield 70%, m.p. 121°C. (from methanol), [α]<sub>D</sub><sup>20</sup> + 9.0 ± 0.5° (c 2.1; CHCl<sub>3</sub>). In the same way as for I, from 2,3,4-triacetyl-β-D-arabinopyranosyl bromide and thiurea 2,3,4-triacetyl-β-D-arabinopyranosyl isothiuronium bromide (VI) was synthesized, m.p. 172°C. (alcohol), [α]<sub>D</sub><sup>20</sup> -26.5 ± 0.8° (c 2.1; alcohol). The boiling of an aqueous solution of NaHSO<sub>3</sub> with VI for 5 minutes and drying of the syrup over sulfuric acid converted it to 2,3,4-triazetyl-β-D-arabinopyranosyl

Card : 4/6

CZECHOSLOVAKIA/Organic Chemistry. Synthetic Organic Chemistry. G-2

Abs Jour: Ref Zhur-Khim , No 24, 1958, 81580.

Author : Hermanek S., Stanek J.

Inst :

Title : A New Method for Obtaining Homoisovaniline.

Orig Pub: Chem. listy, 1958, 52, No 2, 355-56

Abstract: Homoisovaniline (II) was synthesized by the ozonolysis of 4-C<sub>6</sub>H<sub>5</sub>CH<sub>2</sub>O-3-CH<sub>2</sub>OC<sub>6</sub>H<sub>3</sub>CH=CH<sub>2</sub> (I) followed by the hydrogenation of the ozonide formed, and the simultaneous hydrolysis of the C<sub>6</sub>H<sub>5</sub>CH<sub>2</sub>O group. In the same way homoveratrine aldehyde (III) was obtained from the methyl ester of eugenole. The solution of I (15 grams) in ethylacetate (200 ml) was purged with O<sub>3</sub> containing 3% of O<sub>2</sub> (for 6 hours with cooling), afterwards it was hydrogenated for 3 hours on 5%

Card : 1/3

17

STANEK, JAROSLAV

7  
Formation of 2-alkylmercapto derivatives of 4-methyl-6-hydroxypyrimidine Jaroslav Stanek (Karlova Univ., Prague). *Chem. listy* 57, 357-8 (1953).—Suspending 27.8 g.  $\text{HN}:\text{C}(\text{NH}_2)\text{SMe} \cdot \text{H}_2\text{SO}_4$  in  $\text{H}_2\text{O}$ , adding 13 g.  $\text{AcCH}_2\text{CO}_2\text{Et}$  and 10 g. anhyd.  $\text{Na}_2\text{CO}_3$ , keeping the mixt. 4 days with occasional shaking, and recrystg. the insol. condensation product from  $\text{EtOH}$  gave 89.7% 2-methylthio-4-methyl-6-hydroxypyrimidine, m. 218-22°, identified by desulfuration by boiling with Raney Ni in  $\text{EtOH}$  yielding 4-methyl-6-hydroxypyrimidine (I), m. 148-50°. Similarly was obtained 83.4% 2-ethylthio-4-methyl-6-hydroxypyrimidine, m. 145°. Condensation of methylenebis(isothiuronium) bromide with  $\text{AcCH}_2\text{CO}_2\text{Et}$  failed but the 1,2-ethylene- and 1,3-propylene derivs. gave by similar procedure, resp., 34.0% 1,2-bis(4-methyl-6-hydroxy-2-pyrimidylthio)ethane (III), m. 210-20°, and 12% 1,3-bis(4-methyl-6-hydroxy-2-pyrimidylthio)propane, m. 225° (decompos.). Desulfuration of III as above yielded likewise I. Attempts to apply the described method to sugars failed.  
 L. I. Urbánek

4  
 2 May

ja

Country : CZECHOSLOVAKIA G  
 Category : Organic Chemistry. Natural Substances and  
 Their Synthetic Analogs  
 Abs. Jour : Ref Zhur - Khim., No 5, 1959, No. 15491  
 Author : Stanek, J.; Tajmr, L.  
 Institut. : -  
 Title : On Some Derivatives of Quinovose  
 Orig. Pub. : Chem. listy, 1958, 52, No 3, 551-553  
 Abstract : A solution of 20 g. of D-glucose (I) in 300 ml.  
 of pyridine with 20.2 g. of tosyl chloride is  
 left standing for 24 hours at about 20°, 200  
 ml. of pyridine are distilled off, 80 ml. of  
 (CH<sub>3</sub>CO)<sub>2</sub>O are added and after further standing  
 (48 hours at 0°), 1,2,3,4-tetraacetyl-6-p-  
 tosyl-β-D-glucopyranose (II) is obtained, de-  
 composition point 185-190° (from alcohol),  
 [α]<sub>D</sub>+24° (c 0.49; chloroform); the yield of  
 pure substance is 14 g. From the mother liquors,

Card: 1/7

Country : G  
 Category :  
 Abs. Jour : Ref Zhur - Khim., No 5, 1959, No. 15491  
 Author :  
 Institut. :  
 Title :  
 Orig. Pub. :  
 Abstract cont'd. : 5 g. of 1,2,3,4,6-pentaacetyl-β-D-glucopyranose  
 are obtained. By boiling 1,2,3,4-tetraacetyl-  
 6-desoxy-β-D-glucopyranosil-6-isothiouronic  
 acid iodide (Hardegger, E., Montavon, R.M.,  
 Helv. chim. acta, 1946, 29, 1199) with the  
 same weighed quantity of NaHSO<sub>3</sub> in a ten-fold  
 quantity of water in the course of three min-  
 utes, 1,2,3,4-tetraacetyl-6-desoxy-6-mercapto-  
 β-D-glucopyranose (III) is formed, with yield  
 of 68%, m.p. 108°, [α]<sub>D</sub>+9° (c 3.6; chloroform).

Card: 2/7

Category :

Pub. Jour : Ref Zhur - Khim., No 5, 1959, No. 15491

Author :

Institut. :

Title :

Orig. Pub. :

Abstract : (c 0.6; chloroform). Analogously, V is formed  
cont'd. from III, with yield of 87%. 62 g. of tosyl  
chloride are added to 20 g. of I in 300 ml.  
of pyridine, and after the usual treatment,  
60 ml. of  $(CH_3CO)_2O$  is acetylated. While stand-  
ing in a refrigerator, II is separated out.  
After distillation of the mother liquors, 11 g.  
of 1,3,4-triacetyl-2,6,-di-p-tosyl- $\alpha$ ,D-glucopyranose is obtained, m.p. 168° (from  $CH_3OH$ ),  
[ $\alpha$ ]<sub>D</sub> +97° (c 0.45; chloroform), which by in-

Card: 4/7

G - 76

STANSK, J.; CERNY, M.; VRKOC, J.

"Preparation of acylated glycopyranosyl mercaptans." In German. p. 64.

COLLECTION OF CZECHOSLOVAK CHEMICAL COMMUNICATIONS, Praha, Czech.,  
Vol. 24, No. 1, Jan 1959.

Monthly List of East European Accessions (EEAI), LC, Vol. 8, No. 6, Sept. 59  
Unclassified

COUNTRY : CZECHOSLOVAKIA  
 CATEGORY : Organic Chemistry. Synthetic Organic Chemistry  
 REF. JOUR. : RZKhim., No. 1 1960, No.1138  
 AUTHOR : Ernest, I.; Stanek, J.  
 INST. : -  
 TITLE : Decomposition of Diazoketones with Cupric Oxide.  
 V. A New Reaction of Aliphatic Unsaturated  
 $\gamma$ -Diketones  
 ORIG. PUB. : Collect. Czechosl. Chem. Commun, 1959, 24,  
 No 2, 530-535  
 ABSTRACT : no abstract  
 See RZKhim., No 2, 1959, No 4619.

CARD:

1/1

2-6

COUNTRY : CZECHOSLOVAKIA  
CATEGORY : Organic Chemistry. Synthetic Organic Chemistry  
ABS. JOUR. : RZhKhim., No. 23 1959, No. 82249  
AUTHOR : Hermanek, S.; Stanek, J.  
INST. : -  
TITLE : A New Method of Synthesis of Homoisovanillin  
ORIG. PUB. : Collect. Czechosl. Chem. Commun., 1959, 24,  
No 4, 1366-1368  
ABSTRACT : No abstract.  
See RZhKhim., 1958, No 24, No 81580

CARD:

1/1



ERNEST, I.; STANEK, J.

Decomposition of diazoketone with copper (II) oxide. Part 9: Kinetics of cyclization of unsaturated  $\gamma$ -diketone. Coll Cz Chem 26 no.4: 1039-1047 Ap '61.

1. Institut fur organische Chemie, Technische Hochschule fur Chemie, Prag (for Stanek) 2. Forschungsinstitut fur Pharmazie und Biochemie, Prag (for Ernest)

(Diazo compounds) (Copper oxides) (Ketones)

37

1. The purpose of this report is to provide information on the status of the research and development of the various types of weapons and equipment which are being developed for use in the future. This information is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

2. The information in this report is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

3. The information in this report is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

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6. The information in this report is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

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9. The information in this report is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

10. The information in this report is being provided for the use of the various agencies of the Department of Defense and the various agencies of the Department of State.

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STANLEY, J.

STANEK, J

CZECHOSLOVAKIA

MYSKA, J; STANEK, J; ETTTEL, V; MARCHALINOVA, M.

Institute of Organic Technology of the Technical Higher  
School of Chemistry (Institut für organische Technologie,  
Technische Hochschule für Chemie), Prague and Spolana,  
Heratovice

Prague, Collection of Czechoslovak Chemical Communications,  
No 11, 1963, pp 3154-3158

"On Some New Derivatives of Trichloracrylic Acid."

(4)

STANEK, Jaroslav; SMEKAL, Emil

On the possibility of distinguishing chemical compounds by  
semiconducting sensors. Scr. med. fac. med. Brunensis 36  
no.5:215-228 '63.

1. Katedra lekarske fyziky lekarske fakulty university J.E.  
Purkyne v Brne. Vedouci: MUDr. Jaroslav Stanek CSc.  
(CHEMISTRY, ANALYTICAL) (ALCOHOLS)

STANEK, J.

"Organic chemistry. Experiments on a semimacro scale" by  
G.H. Coleman, S. Wawzonek, R.E. Buckles. Reviewed by J. Stanek.  
Chem listy 57 no.6:662-663 Je '63.

STANEK, J.

"Structural carbohydrate chemistry" by E.G.V. Percival,  
E. Percival. Reviewed by J. Stanek. Chem listy 57 no.11:  
1201-1203 N '63.

STANEK, J.; SINDLEROVA, M.; CERNY, M.

Derivatives of D-thioxylopyranose and of some reducing 1-deoxy-1-thiosaccharides. Coll Cz Chem 30 no.1:297-303 Ja '65.

1. Department of Organic Chemistry of Charles University, Prague. Submitted April 30, 1964.

STANEC, Jaroslav, Ing.

The 11-class aggregate of 200 mva transformers of the AEG electric  
power plant for 380 kv network in the German Federal Republic. El  
tech. doc. 53 no.9:504-106 S '64

1. Institute of Electrical Engineering, Czechoslovak Academy of  
Sciences.



STANEK, J.

Nomenclatural rules of organic chemistry. Chem listy 59 no.3:265-327 Mr '65.

SENFT, J.

Problem in the production of light bottles.

p. 342 (Sklar A Keramik. Vol. 7, no. 11, Nov. 1957. Praha, Czechoslovakia)

Monthly Index of East European Accessions (EEAI) LC. Vol. 7, no. 2,  
February 1958

STANEK, J.

"Trends of technical development in the glass industry." (To be contd.)  
P. 99.

SKLAR A KERAMIK. (Ministerstvo lehkeho prumyslu). Praha, Czechoslovakia,  
Vol. 9, No. 4, Apr. 1959.

Monthly list of East European Accessions (EEAI), LC, Vol. 8, No. 8,  
August 1959.  
Uncla.

COUNTRY:	: Czechoslovakia	H-15
CATEGORY	:	18757
ABST. JOUR.	: RZKhim., No. 5 1960, No.	
AUTHOR	: Stanek, J.	
TITLE	: Not given	
TECH	: New Technological Trends in the Glass Industry	
ORIG. PUB.	: Sklar a Keramik, 9, No 5, 131-134 (1959)	
ABSTRACT	: A number of the properties of glasses are listed: electric conductivity and its relationship to chemical composition and viscosity of the glass, dependence of the electric resistance on the temperature of the material, and a calculation of the energy required for the electric melting of the glass. The technology of electric melting in small units and in so-called 'deep' electric furnaces is described. In furnaces of recent design the electrodes are placed in such a way that the lines of	

COUNTRY : Czechoslovakia  
CATEGORY :

n-12

INST. JOUR. : RZKhim., No. 5 1960, No.

16757

AUTHOR :  
INST. :  
TITLE :

ORIG. PUB. :

ABSTRACT

: force pass uniformly and simultaneously through the entire cross section of the glass. Large-area electrodes are used with a resulting decrease in the current density at the contact with the refractory and an increase in the service life of the latter; this arrangement simultaneously eliminates the formation of gas bubbles in the glass mass. The construction of a furnace of the above design requires about 9 times less refractory material than that of furnaces of standard design. The pro-

271

STANEK, Jaroslav; GOTZ, Jiri

Examination of the microstructure of thermally refined glass surface by means of electron microscope. Silikaty 6 no.3:291-295 '62.

1. Vysoka skola strojni a textilni, Liberec; Vyzkumny ustav uzitkoveho skla a bizuterie, Jablonec nad Nisou.

STANEK, J.

"Trends of technical development in the glass industry." P. 164.

SKLAR A KERAMIK. (Ministerstvo lehkeho prumyslu). Praha, Czechoslovakia,  
Vol. 9, No. 6, June 1959.

Monthly list of East European Accessions (EEAI), LC, Vol. 8, No. 8,  
August 1959.  
Uncla.

GOTZ, Jiri; STANEK, Jaroslav

Effect of mechanical treatment on the microheterogenous structure of glass surfaces. Silikaty 7 no.1:42-45 '63.

1. Vyzkumny ustav uzitkoveho skla a bizuterie, Jablonec nad Nisou (for Gotz). 2. Vysoka skola strojni a textilni, Liberec (for Stanek).



STANEK, Jaroslav

"Glass technology." Reviewed by Jaroslav Stanek. Sklar a keramik 13  
no.2:55 F '63.

JOSZT, Kazimierz; KRZYZANOWSKI, Wojciech; STANEK, Jerzy

Testing oxidation of M70 and M63 brasses. Mechanika Gliwice  
no.16:97-110 '62.

1. Katedra Metaloznawstwa, Politechnika, i Instytut Metali  
Niezelaznych, Gliwice.

STANEK, Jindrich

621.314.2  
4908. The 220/110 kV transformers of the CKD  
Stalingrad National Corporation. V. VOCHOČ AND  
J. STANEK. *Elektrotech. Obzor*, 42, No. 9, 504-10  
~~1993-78-678ch~~

Detailed description of the design, construction and

testing of the standard 100 MVA 220  $\pm$  9  $\times$  3  $\cdot$  3/110/  
10.4 kV transformer of the CKD works. This  
transformer is designed as three 1-ph. units with one  
unit spare and the voltage regulation is being done  
by a separate regulating transformer connected to the  
neutral end of the 220 kV winding. H. NOREL

STANEK, J. (Jindrich)

621.317.7

683. CONSIDERATION OF SOME PROBLEMS IN ELECTRICAL MEASUREMENTS. J. Stanek.

Nachrichtentechnik, Vol. 7, No. 9, 423-32 (Sept., 1957). In German.

The measuring techniques and apparatus described include:

photoelectric amplifiers using the self-balancing electronic method for measurement of direct currents and voltages, mechanical forces and torques; a polarograph using a p.-e. amplifier; a mains-energized portable radiation dose-meter using a quadrant electrometer across a  $10^8$ - $10^9$  MΩ resistor; a pocket dose-meter using an electrometer and based on the capacitor loss-of-charge method; a thermal convertor for measurement of r.f. currents which uses a hot diaphragm and thermocouples instead of the usual hot wire and thermocouple, the resulting reduction of skin effects extending the frequency range.

C. F. Pizzet

STANEK, JINDRICH

✓ 7440\* (Czech.) The Technology of Electro-Spark Machining.  
Technologie elektroiskrového obrábění. Jindřich Stanek.  
Elektrotechnik, v. 12, Jan. 1957, p. 2-7.

Method for machining hard materials based on the erosion of  
the metal by spark discharges between the tool electrode and  
the part being machined.

STANEK, J.

"Electric spark machining of metals."

p. 33 (Czechoslovak Heavy Industry, no. 5, 1958, Praha, Czechoslovakia)

Monthly Index of East European Accessions (EEAI) LC, Vol. 7, no. 9,  
September 1958

STANEK, J.

TECHNOLOGY

ELEKTROTECHNICKY. OBZOR.

STANEK, J. Resolution of magnetic fluxes in transformer regulating sets for locomotives. p. 616.

Vol. 47, no. 12, Dec. 1958.

Monthly List of East European Accessions (EEAI), LC, Vol. 8, no. 5  
May 1959, Unclass.

STANEK, Jindrich

"Electric spark machining" by Andras Roth, Ferenc Kaldos  
and Istvan Kovacs. Reviewed by Jindrich Stanek. Stroj  
vyr 10 no.10:535 0 '62.



STANEK, Jindrich, inz.

"Ultrasound machining" by [doc., dr.] Jerzy Dmochowski. Reviewed by  
Jindrich Stanek. Stroj vyr 10 no.12:635 '62.

STANEK, Jindrich, inz. CSc.

Problems of the mechanization and automation of assembling.  
Stroj vyr 11 no.11:542-546 N°63.

1. Riaditel, Vyvojovy ustav pre mechanizaciu a automatizaciu, Nove Mesto nad Vahom.

ACC NR: AP6035580

SOURCE CODE: CZ/0047/66/000/010/0835/0838

AUTHOR: Stanek, Jindrich (Candidate of sciences; Engineer)

ORG: none

TITLE: Academic community near Novosibirsk

SOURCE: Technicka praca, no. 10, 1966, 835-838

TOPIC TAGS: academic institution, academic personnel, scientific program, scientific personnel / Akadem-Gorodok

ABSTRACT: The Siberian Branch of the Academy of Sciences, USSR has its scientific and administrative center in the Academic community (Akadem-Gorodok) near Novosibirsk, located on the shore of the "Ob' Sea," a water reservoir of about 6,000 km behind the dam. In this nine-year old community of 35,000 persons, there are 16 institutes of the Academy with more than 1200 scientific workers, among which are 10 Academicians, 30 Corresponding Members, 100 Doctors of Sciences, and 1100 Candidates of Sciences. The chairman of the Branch, since its foundation in 1957, is Academician M. A. Lavrent'yev. The following objectives are pursued by the Branch with a tendency to combine science with experience: 1) To establish a large collective of scientific workers in Siberia to solve significant problems encountered in the development of Siberian industry; 2) To educate scientific personnel by practical training; 3) To promote the development of the national economy in Siberia and the Far East. It is

Card 1/2

ACC NR: AP6035580

planned to have here in 1970, twelve large design bureaus and institutes for applied research, experimental factories for further development of basic-research discoveries, a polytechnic institute, and trade schools. Some working models (of a new particle accelerator, a hydrodynamic device for super-high pressure impulses) produced in the community were shown to the author. The methods of selecting future scientists from schoolboys and their education, which starts in high school and is continued at the university which has 3000 students and 500 candidates, are described, as well as the organization and operation of the experimental institutes. Novosibirsk itself is a city of more than one million, has 14 schools of higher education, and more than 50 scientific research institutes.

SUB CODE: 05 / SUBM DATE: none

Card 2/2

Distr: hE2c(j)/hE3d

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4  
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2

Estimation of pentaerythritol tetranitrate in the presence of cyclotrimethylene trinitroamine. <sup>7</sup> J. Stanek and Jan Vacek. *Chem. průmysl* 8(33), 361-3(1958).—Pentaerythritol tetranitrate (I) and cyclotrimethylene trinitroamine (II) can be detd. simultaneously by the following method: Dissolve a sample contg. not more than 0.22 g. of I and 0.12 g. of II in 30 ml. of glacial AcOH and add 25 ml. of 25% HCl. Remove air by passing a stream of CO<sub>2</sub> through the soln. Add 50 ml. of 0.7N FeCl<sub>3</sub> and reflux for 20 min. Cool, add 10 ml. of 10% NH<sub>4</sub>CNS, and titrate with 0.2N TiCl<sub>3</sub> until the red color disappears. Then add 20 ml. of N TiCl<sub>3</sub> and 25 ml. of Na citrate and reflux for 30 min. Cool and titrate with 25 ml. of 0.25N FeNH<sub>4</sub>(SO<sub>4</sub>)<sub>2</sub> until the red color persists for 1 min. Blanks have to be detd. and their values subtracted. The amt. of I is equiv. to the cor. vol. of TiCl<sub>3</sub>, that of II to the cor. vol. of FeNH<sub>4</sub>(SO<sub>4</sub>)<sub>2</sub>. Max Hellmann,

Jed

STANEK, J.; DRAHONOVSKY, J.; VACEK, Z.

On the mechanism of the biological action of some nitrophenols. Folia  
microbiol. 8 no.1:48-55 '63.

(NITROPHENOLS)

(NEOROSPORA)

(HYDROGEN ION CONCENTRATION)

VACEK, Z.; STANEK, J.

Chromatography of chloronitrophenols. Coll Cz Chem 28 no.1:264-266  
Ja '63.

1. Spolana, Neratovice.

MANDIK, Lomir; STANEK, Jiri

Selection of solvents for coating systems according to the solubility parameter. Chem prum 15 no.4:223-226 Ap '65.

1. Research Institute of Synthetic Resins and Lacquers, P. dubice..



KASIK, Ivan, inz.; STORKAN, Zdenek, inz.; STANEK, Jiri

Electroslag remelting of steel and alloys in the Soviet  
Union. Hut listy 19 no.5:311-318 My '64

1. Research Institute of Iron Metallurgy, Prague (for Kasik  
and Stanek). 2. Spojene ocelarny, Kladno (for Storkan).

1A STANEK, JOSEF

*Mineralogical*

Beryl pegmatite of Scholbengraben near Martitz, Moravia. J. Pokorný and J. Staněk (Masaryk Univ., Brno, Czech.). *Práce Moravskoslezského ústavu geologického* 23, 247-68 (1951) (English summary).—The pegmatite, intrusive into amphibolitic gneiss, is complex. The earliest phase contained K feldspars, quartz, tourmaline, mica, beryl, and columbite; a later albite-rich phase contains quartz, mica, garnet, spinel, and beryl. The last minerals to form in cavities were tourmaline, apatite, albite, mica, cuculite, bertrandite, and quartz. The pegmatite is too lean to be worked commercially. Michael Fleischer

STANEK, J.

# CZECH

Cyrllovit, a new phosphate mineral. M. Novotný and J. Štáhl (Masaryk Univ., Brno, Czech.). *Průmyslová geologie* 10, 1963, 126-30 (1963) (English summary).  
 The mineral occurs as rare crystals with triplite, triplidite, and dufrénite in nests in pegmatite at Cyrllov, West Moravia. It occurs as brown tabular tetragonal crystals with  $a:b:c = 1:0.5314$  (goniometric). Rotation photographs confirm the tetragonal symmetry, but lead to a different orientation with  $a: b: c = 7.32: 1: 19.4$  Å. Analysis gave  $P_2O_5$  33.06,  $Fe_2O_3$  50.50,  $H_2O$  15.14, sum 98.69%. The mineral is uniaxial neg.,  $n_O$  1.805,  $n_E$  1.777. Sp. gr. is 3.085. X-ray powder data are given. Michael Fleischer.

STANEK J.

/ Occurrence of molybdenite in the plutonic rocks near  
Zulová (Czechoslovakia). J. Staněk and J. Kolář.  
Věstník Ústřed. Ústavu Geol. 28, 271-7(1953).—Molybde-  
nite is present only in small amts. in veins of pegmatite.

Other accessory minerals found included fergusonite, phen-  
akite, and monazite. GP.  
Rudolph Pick

(2)

STANEK, J.

# CZECH

STANEK, J.

✓ Mineralogy of some West Moravian pegmatites. J. Staněk (Univ. Brno, Czech.). *Casopis Moravského Muzea* 38, 112-29(1953)(English summary).—S. identified from pegmatites near Dolní Bory scorodite, loellingite, pharmacosiderite (ns 1.697, 1.704), symplectite with ns  $\alpha$  1.633,  $\beta$  1.646,  $\gamma$  1.681, augelite with ns  $\alpha$  1.573,  $\gamma$  1.589, sp. gr. 2.687, and metamict zircon high in P, Y, and Ca content. Amblygonite with ns (Na)  $\alpha$  1.590,  $\gamma$  1.610, sp. gr. 3.11, and crystals of columbite were found at Foklce. Beryl of 2 types (ns  $\alpha$  1.575, 1.592,  $\epsilon$  1.570, 1.578), topaz with ns  $\alpha$  1.618,  $\beta$  1.621, bertranidite with ns  $\alpha$  1.589,  $\gamma$  1.610, and petalite with ns (Na)  $\alpha$  1.506,  $\beta$  1.512,  $\gamma$  1.518, sp. gr. 2.393, were found near Jeclov. Michael Fleischer

STANEK, JOSEF

CZECH

Spodumene and bavenite from Jeclov. Josef Stanek.  
Casopis moravského Muzea, Vědy přírod 39, 82-77 (1954).  
(English summary).—Spodumene occurs as an alteration  
product of petalite, which also alters further to a montmoril-  
lonitic clay. Spodumene had  $n_s \alpha 1.658$ ,  $\gamma 1.674$ . Bavenite  
occurs in cavities assocd. with adularia. It had  $n_s \alpha 1.587$ ,  
 $\beta 1.589$ ,  $\gamma 1.594$ . Michael Fleischer

dc

STANEK, JOSEF

3

✓ The phosphate paragenesis from Cyrilov, West Moravia.  
Josef Staněk. *Časopis Moravského Muzea, Vědy Přírod*  
~~40, 66-68 (1955)~~ (English summary).—Primary minerals in-  
clude graptolite with  $ns \alpha$  1.718,  $\beta$  1.723,  $\gamma$  1.745; apatite  
with  $ns \alpha$  1.635,  $\epsilon$  1.633; manganous apatite with  $ns \alpha$   
1.652,  $\epsilon$  1.648; and triplite with  $ns \alpha$  1.695,  $\gamma$  1.705. Oxi-  
dation products include cyrilovite (cf. C.A. 49, 7454d),  
heterosite, strengite with  $ns \alpha$  1.730,  $\beta$  1.733,  $\gamma$  1.763, sp.  
gr. 2.839; metastrengite with  $ns \alpha$  1.692,  $\gamma$  1.737; and  
fairfieldite (?) with  $ns \alpha$  1.637,  $\gamma$  1.656. Michael Fleischer

CZECHOSLOVAKIA / Cosmochemistry, Geochemistry. Hydrochemistry. D

Abs Jour : Ref Zhur - Khim., 1958, No 17, No 57056.

Author : Stanek, J

Inst : Not given

Title : Bismuth and Bismutite from Pegmatites Near Marsikova in Northern Moravia.

Orig Pub : Casop. Moravskeho musea Brnl, Vedy prirod, 1957, 42, 33-38.

Abstract : Block (quartzitic) and metasomatic (albite) pegmatites are characterized by the following association of minerals: apatite, babenite, hematite, beryl, bertrandite, biotite, bismutite, native bismuth, garnet, columbite, microcline, muscovite, topaz, tourmaline (schorlite), zircon, spinel and euclase. The results of spectrum analyses (in %): native bismuth  $\geq$  IBi, N.  $10^{-2}$  Ca and Si,  $\leq$  0.01, Ag, Al, Cr, Cu, Fe, Mg, Mn, Pb and Sb; bismutite -  $\geq$  IBi, 0.01 Al,

Card 1/2



STANEK, J.

GEOGRAPHY & GEOLOGY

Periodicals: CASOPIS PRO MINERALOGII A GEOLOGII Vol. 3, no. 4, 1958

CECH, F.; STANEK, J. A new lithium pegmatite from Nova Ves near  
Cesky Krumlov. p. 407.

Monthly List of East European Accessions (EEAI) LC, Vol. 8, No. 5  
May 1959, Unclass.

STANEK, Josef; MISKOVEKY, Josef

Cordierite rich in iron from the pegmatite deposit near Jolín  
Bory. Cas min geol 9 no.24191-192 '64.

1. Faculty of Natural Sciences, J.E. Purkyně University, Brno;  
Higher School of Technology, Brno.

STANEK, Karel; VAVERKA, Miroslav, inz.

Experience with the construction of the press shop in the  
enterprise Zavody na vyrobu kulickovych lozisek. Poz stavby  
II no.1:24-26 '63.

1. Ingstav, Brno.

STANER, L.; HYKEL, J.

Detonators, primers and electric detonators. p. 374. (STAVIVO, Vol. 35, No. 9, Sept 1957, Praha, Czechoslovakia)

CO: Monthly List of East European Accessions (SEAL) LC, Vol. 6, No. 12, Dec 1957. Uncl.

STANEK, Liboslav, inz.

Construction of thermal power stations using lignite. Energetika  
Cz 11 no.9:425-426 S '61.

STANEK, Liboslav, inz.

To reach the world standard in construction of condensing electric power plants and in distribution of electric power. Energetika 12 no.1:Suppl.:1-5 Ja '62.

1. ~~Minister~~stek ministra paliv a energetiky.

STANEK, L., inz.

Entering the second year of the third Five-Year Plan. Paliva  
42 no.1:1-2 Ja '62.

1. Namestek ministra paliv a energetiky, Praha.

STANEK, Liboslav, inz.

Third International Mining Congress in Salzburg.  
Uhli 5 no.11:386-387 N '63.

1. Clen mezinarodniho hornickeho komitetu, Praha.



STANEK, Liboslav, inz.

Science and technology, the most efficient productive forces.  
Uhli 6 no. 2:41-43 F '64.

1. Glen kolegia ministra paliv.

STANEK, Liboslav, <sup>1</sup>prz.

Drilling and blasting methods and their importance for the development of the mining of minerals. Rudy 12 no.2: 41-42 F'64

1. Statni komise pro rozvoj a koordinaci vedy a techniky, Praha.

STANEK, Liboslav, inz.

Planning of the science and research in the mining and deposit survey. Rudy 12 no. 3: 73-76 Mr '64.

1. State Commission for the Development and Coordination of Science and Technology, Prague.

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**APPROVED FOR RELEASE: 08/25/2000** **CIA-RDP86-00513R001652820011-5"**

"Preparation for the second year of industrial high schools of mining" by [inz.] Z. Masin, [inz.] A. Kostal.  
Reviewed by Liboslav Stanek. Uhli 6 no. 4: 142 Ap '64.

"Preparation of mineral raw materials" by [inz.] Oskar Dinter. Reviewed by Liboslav Stanek. Ibid.: 142.

STANEK, Liboslav, inz.

Rational use of domestic fuel and power resources. Tech  
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1. State Commission for the Development and Coordination  
of Science and Technology, Prague.

STANEK, L., inz.

Remarks on the M.F. Kessler article "Main trends in coal use and research." Paliva 44 no.10:319-320 O '64.

1. State Commission for Development and Coordination of Science and Technology, Prague.

STANEK, I.

Development of longwall mining in the North-Bohemian Lignite Basin. p. 15.

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Praha, Czechoslovakia  
Vol. 1, no. 1, Jan. 1959.

Monthly list of East European Accessions (EEAI), LC, Vol. 8, no. 7  
July 1959  
Uncl.

STANEK, L.

Data on Soviet coal mines and fuels. p. 209.

UHLI. (Ministerstvo paliv) Praha, Czechoslovakia. Vol. 1, no. 6, June 1959

Monthly list of East European Accessions (EEAI), Vol. 9, no. 1, Jan. 1960

Uncl.

STANEK, L.

Saving our forests and saving timber for mines. p. 329.

UHLE. (Ministerstvo paliv) Praha, Czechoslovakia,  
Vol. 1, no. 10, Oct. 1959.

Monthly List of East European Accession (EEAI), LC Vol. 9, no. 2,  
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Uncl.



STANEK, L.

Further rise in the productivity of brown coal mines of Czechoslovakia. p. 363

UHLI (Ministerstvo paliv) Praha, Czechoslovakia. Vol. 1, no. 11, Nov. 1959

Monthly list of East European Accessions (EEAI), Vol. 9, no. 1, Jan. 1960

Uncl.

STANEK, L.

Preparation of coal of the north Bohemian area for power purposes. p. 269.

PALIVA. (Ministerstvo paliv a Československá vědecká technická společnost pro využití paliv při Československé akademii věd) Praha, Czechoslovakia, Vol. 39, no. 8, August 1959.

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uncl.

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Determining basic indexes for evaluation of open mining conditions. Uhli 5 no.5:167-170 My '63.

1. Statni komise pro rozvoj koodinace vedy a techniky,  
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STANEK, L., inz.

Mechanization of low seam mining. Uhl 6 no. 8:287 Ag '64.

STANEK, K.

"Visibility and Flying", P. 328, (KRIDLA VLASTI, Vol. 4, No. 14,  
July 1954, Praha, Czechoslovakia)

SO: Monthly List of East European Accessions, (EEAL), LC, Vol. 4,  
No. 1, Jan. 1955, Uncl.

STANEK, M.

Glider, clouds, and breathing apparatus. p.230.  
KRIDL VLASTE, Prague, No. 10, May 1955.

SO: Monthly List of East European Accessions, (EEAL), LC, Vol. 5, No. 6 June 1956, Uncl.

KLAZAR, J.; KOSEK, F.; CIMPL, J.; STANEK, M.

Color temperature method for determining the heating of  
parts. Automatizace 7 no. 3:81 Mr '64.

STANEK, M., dr.; UJEVIC, I., inz.; NOVAKOVA, J., inz.; DOLEZALOVA, H.,  
inz.

Effect of the Indian corn seed treatment by preparations containing tetramethylthiuram disulfide on the development of parasitic, saprophytic, and mycolytic microflora around the planted seed. Rost vyroba 10 no. 4:423-434 Ap '64.

1. Institute of Microbiology, Czechoslovak Academy of Sciences, Prague (for Stanek). 2. Central Research Institute of Plant Production, Ruzyně (for Ujevic). 3. Central Control and Testing Institute of Agriculture, Brno (for Novakova).



CZECHOSLOVAKIA/Electronics - Photocells and Semiconductor Device H.

Abs Jour : Ref Zhur - Fizika, No 7, 1959, 15961

Author : Stanek, M., Medvezov, B.

Inst : Institute for Radio Technical Research, Czechoslovakia

Title : Meter for the Input Parameters of Transistors

Orig Pub : Slaboproudny obzor, 1958, 19, No 3, 128-132

Abstract : Description of a device that makes possible the input parameters of transistors (input impedances and input capacitance) by means of a compensation method with the aid of a rectangular signal. From the shape of the different signal on the oscillograph screen it is possible to determine which of the parameters of the electric equivalent circuit does not match the corresponding input parameter of the measured transistor.

Card 1/1

43034

S/194/62/000/010/074/084  
A055/A126

AUTHOR: Staněk, Milan

TITLE: Selective amplifier system

PERIODICAL: Referativnyy zhurnal, Avtomatika i radioelektronika, no. 10, 1962,  
102, abstract 10-7-204b P (Czech. pat., cl. 21a<sup>4</sup>, 29/02, no. 98244,  
January 15, 1961)

TEXT: The patented system (see Figure) contains a summation element S to whose input is applied the amplified signal and from whose output is picked up the voltage applied to the amplifier A that converts the phase of the amplified voltage. The output of the amplifier (which is also the output of the whole system) is connected to the input of the integrating circuit I and to one of the inputs of the comparison circuit P, whose second input is connected to the output of I. If the voltage at the output of P in absolute value deviates from the voltage at the input of I, the input impedance of P will vary and, as a result, the level of the signal at the input of this circuit will again become equal in absolute value, and be phase-shifted by 90° with respect to the voltage at the in-

Card 1/2

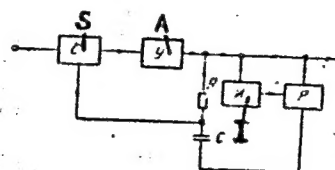
# Selective amplifier system

S/194/62/000/010/074/084  
A055/A126

put of I. These voltages are applied to the series circuit RC, and the equality  $f = \frac{1}{2\pi RC}$  is valid for them, f being the transmission frequency. The advantage of the system is that it permits the control of f within wide limits by means of the variation of one of the elements R, C, or the attenuation of the cascade N - P [Abstracter's note: N seems to be a misprint for H = 1]. A practical transistorized variant of the proposed device is described.

H.P.

[Abstracter's note: Complete translation]



Card 2/2

41252

S/194/62/000/007/146/160  
D413/D308

9.1.17  
AUTHOR: Staněk, Milan

TITLE: A DC-connected transistor amplifier

PERIODICAL: Referativnyy zhurnal. Avtomatika i radioelektronika,  
no. 7, 1962, abstract 7-7-227 i. (Czech. pat., cl. 21a<sup>2</sup>  
18/08, no. 96931, Oct. 15, 1960)

TEXT: The patent describes the circuit and selection of parameters for the amplifier. The circuit consists of one resistor, one capacitor and four transistors: two p-n-p and two n-p-n type. The output is push-pull (paraphase), and the current amplification factor is about 60 dB. The recommended form of load is a center-tapped output transformer, remarkable for its simplicity of construction, economy, small dimensions and high stability, which is achieved by using an auxiliary circuit containing one transistor whose base is fed with a fixed boost voltage equal to half the supply voltage. The possibility is discussed of varying the parameters of the circuit for practical applications. The basic advantages of the circuit are its high stability and the absence of an input transformer  
Card 1/2